

## CHEMICAL OXYGEN DEMAND REDUCTION FROM COFFEE PROCESSING WASTE WATER – A COMPARATIVE STUDY ON USAGE OF BIOSORBENTS PREPARED FROM AGRICULTURAL WASTES

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### ABSTRACT

The utilization of three novel and distinct agricultural waste materials, namely casuarina fruit powder (CFP), sorghum stem powder (SSP) and banana stem powder (BSP) as low cost adsorbents for reducing chemical oxygen demand (COD) levels of coffee industry wastewater has been examined. Prepared adsorbents were characterized by scanning electron microscopy (SEM), Energy-dispersive x-ray spectroscopic (EDAX) and Fourier Transform Infrared Spectroscopy. The effects of parameters like adsorbent dose, pH, temperature, agitation rate and time of adsorption on the reduction of COD levels were analyzed for each adsorbent individually and the efficiency order is CFP>SSP>BSP. Based on the extent of compatibility to Freundlich/Langmuir adsorption isotherm and different models (pseudo-first order and second order, Boyd, Weber's and Elovich), simultaneous occurrence of chemisorption and physisorption was proposed. Based on the observations, it is proposed that three kinetic stages involve in adsorption process viz., diffusion of sorbate to sorbent, intra particle diffusion and then establishment of equilibrium. As these three adsorbents exhibit adsorption capacity more than 80%, the adsorbents prepared from agricultural wastes have promising role towards reduction of COD from industrial wastewater to contribute environmental protection.

**Keywords:** Coffee industry waste water, COD, adsorption, casuarinas fruit, sorghum stem, banana stem

### 1. Introduction

Coffee is ranked second among traded commodities, where India has a share of 4.5% in global market and dictates the export economy. But, coffee marketing is facing environmental and social issues (Tallontire, 2002). In wet coffee processing procedure, removal of pulp (de-pulping) consumes more than half of the water required for entire process (Fresner and Schnitzer, 1996). One tonne of processed coffee results in pollution equivalent to 45 kg COD or 273m<sup>3</sup> of crude domestic sewage (Calzada *et al.*, 1989). Coffee processing plants generate waste water having colour and different organic molecules like pectin, fibers, fat, caffeine, polyphenols, proteins, carbohydrates (reducible and non-reducible), chlorogenic acid, and other macromolecules (lignins, tannins, humic acid) (Von Enden, 2002) which are rapidly biodegradable. Coffee processing plant effluents cause bad odor and pollution of ground water (Mburu *et al.*, 1991).

Depletion of dissolved oxygen and a decrease in pH (due to fermentation of organic matter) may hamper the sustainability of water bodies (Fresner and Schnitzer, 1996). Studies carried out by Alemayehu Haddis and Rani Devi, (2008) reveals that the consumption of water polluted with coffee processing plant causes diseases like breathing problem, nausea, stomach problem and skin irritation. The reduction of these COD levels plays an important role in the reuse of the industrial effluents (Selvamurugan *et al.*, 2010).

As per WHO (1995) norms, COD permissible limits for irrigation is 300 mg l<sup>-1</sup>. Traditionally, the reduction of COD levels from industrial effluents was achieved by various biological and physico chemical techniques like fermentation, sedimentation, flocculation, coagulation etc. Hence, the minimization of COD to permissible levels is essential for the safe survival of ecosystem. Among all the reported methods, biosorption is considered as a most promising technique because it is rapid, economic and environmentally benign technology and also is not associated with the generation of secondary wastes which present treatment problems (Verma and Mishra, 2010; Ramesh *et al.*, 2014). In several previous reports, many investigators have studied the use of various plant materials as adsorbents for the reduction of COD from waste water, such as, avocado peel or seed carbon (Rani Devi *et al.*, 2008; Rani Devi, 2010), metakolin, dates nut carbon and tamarind nut carbon (Parande *et al.*, 2009). Hence, in this study an attempt is made to develop low-cost, effective and commercially applicable adsorbents from agricultural wastes (casuarina fruit, sorghum stem and banana) and to study their usage for minimizing the COD levels from coffee waste water to cater the needs of developing and underdeveloped countries.

## 2. Experimental

### 2.1 Materials and Methods

The three adsorbents used for this study were (a) Casuarina fruit powder (CFP) (b) Sorghum stem powder (SSP) and (c) Banana Stem powder (BSP). The cleaned raw materials were washed with formaldehyde solution (2%) to reduce organic leaching and prevent mould growth during batch sorption (Chen and Yang, 2005) and completely air dried. Sieves of particles (size 0.3mm to 1.0mm) were used. All the chemicals used were of analytical grade. The adsorption studies were done with washings (sample was collected prior to final stage of effluent treatment in industry and used without any further processing/treatment after collection i.e., 11,300 mg l<sup>-1</sup> as Initial COD) from Continental Coffee Limited, Duggirala, Andhra Pradesh, India. The COD of the waste water samples before and after treatment with adsorbents were measured according to the methods prescribed by the ASTM (American society for testing and materials, 1972) annual book. The pH was adjusted using 0.1M HCl or 0.1M NaOH.

### 2.2 Adsorption studies

The experiments were performed in batch process in Erlenmeyer flask containing 250 ml of wastewater sample and the total volume of the reaction mixture were kept at 100 ml. The equilibrium time for adsorption was determined as 3hours and the other experiments were conducted in batch with common parameters: 37±2 °C temperature, 4 g l<sup>-1</sup> adsorbent, neutral pH, and 200 rpm. The removal efficiency of each adsorbent with respect to different parameters like time (30–240 min), adsorbent dose (100–1000 mg/100 ml), pH (3–10), temperature (30–70 °C), and speed of rotation (50–300 rpm) was determined in individual experiments. In each experiment, the wastewater samples equilibrated with adsorbents were withdrawn, centrifuged and the residual COD concentrations were measured. The experiment was repeated thrice and average results were documented.

## 3. Results and Discussion

### 3.1 SEM, EDAX and FT-IR studies

Properties of the solute and sorbent dictate the nature of adsorption. Hence, the adsorbent characterization was done by scanning electron microscopy (Fig. 1) which visualizes the surface texture and porosity of the samples. Rough surface with lot of holes were visible in form of dark area for

micrographs before adsorption which is an indication of large surface areas, whereas, a smooth surface with some form of cluster having ice coating can be observed after adsorption. Prior to removal of COD, the CFP sample surface consists of cavities indicating high possibility of adsorption. EDAX studies were carried out to determine the chemical composition of biosorbents (Wt% of elements). BSP:C-62.65, O-34.44, Si-1.15, Fe-0.87; CFP:C-46.84, O-40.49, Na-10.71, Si-1.51, Fe-0.45; SSP:C-61.26, O-34.69, Na-0.35, Si-0.25, Fe-0.22, Zn-0.26. In all the adsorbents, presence of stretching for C-O, C=O and aliphatic C-H can be confirmed from peaks in the range of 1023–1039  $\text{cm}^{-1}$ , 1614–1689  $\text{cm}^{-1}$  and 2917–2920  $\text{cm}^{-1}$  respectively. But, additional peaks in CFP above 3000  $\text{cm}^{-1}$  indicate the =C-H stretching and further C=C stretching can be observed from peaks in the range of 1450–1550  $\text{cm}^{-1}$ . Broad peaks in the range of 3305–3320  $\text{cm}^{-1}$  indicate the presence of carboxylic acid groups in SSP and BSP, whereas, dominant sharp peak at 3737  $\text{cm}^{-1}$  indicates phenolic groups in CFP (Fig. 2).

EDAX and FT-IR observations correlate with the compositions of SSP (cellulose, hemicelluloses, lignin, cutin, silica – Rafia and Gilani, 2001), BSP (holocellulose, cellulose, lignin, pectin – Li *et al.*, 2010) and CFP ( $\alpha$ -pinene, benzaldehyde, 1,8 cineole, furanoid,  $\alpha$ -campholenal, 4-terpineol,  $\alpha$ -terpineol,  $\alpha$ -terpinyl acetate, spathulenol, caryophyllene-oxide, guaiol – Ogunwande *et al.*, 2011). Presence of more number of peaks in CFP compared to SSP and BSP can be explained based on presence of good number of phytochemicals as mentioned above. The presence of functional groups like acid, alcohol and amine are evident from these studies. Protons of phenols/alcohols and carboxylic acid groups in these biosorbents facilitate them for participation in ion exchange. The physico-chemical parameters like TSS, TDS, pH etc were determined for the coffee processing waste water and were given in Table 1.

**Table 1.** Initial physico-chemical parameters of coffee processing waste water

| Parameter                           | COD         | N   | Fluoride | Phenolics | pH      | TSS        |
|-------------------------------------|-------------|-----|----------|-----------|---------|------------|
| Concentration (mg l <sup>-1</sup> ) | 11300 ± 208 | 4.0 | Nil      | Nil       | 3.9-4.1 | 3190 ± 355 |

### 3.2 Effect of contact time

The percent of removal increased with an increase in contact time and reached maximum at 180 minutes by attaining equilibrium conditions (Fig. 3) and afterwards, sluggishness was observed which can be attributed to the saturation of binding sites. No decrease in percent of removal after optimum conditions showed the absence of desorption. Though, Muftah H. El-Naas *et al.*, 2010 reported the maximum COD reduction in 30 min, the adsorbent dose used by them was 20 g l<sup>-1</sup>, whereas, it is 4 g l<sup>-1</sup> in the present case.

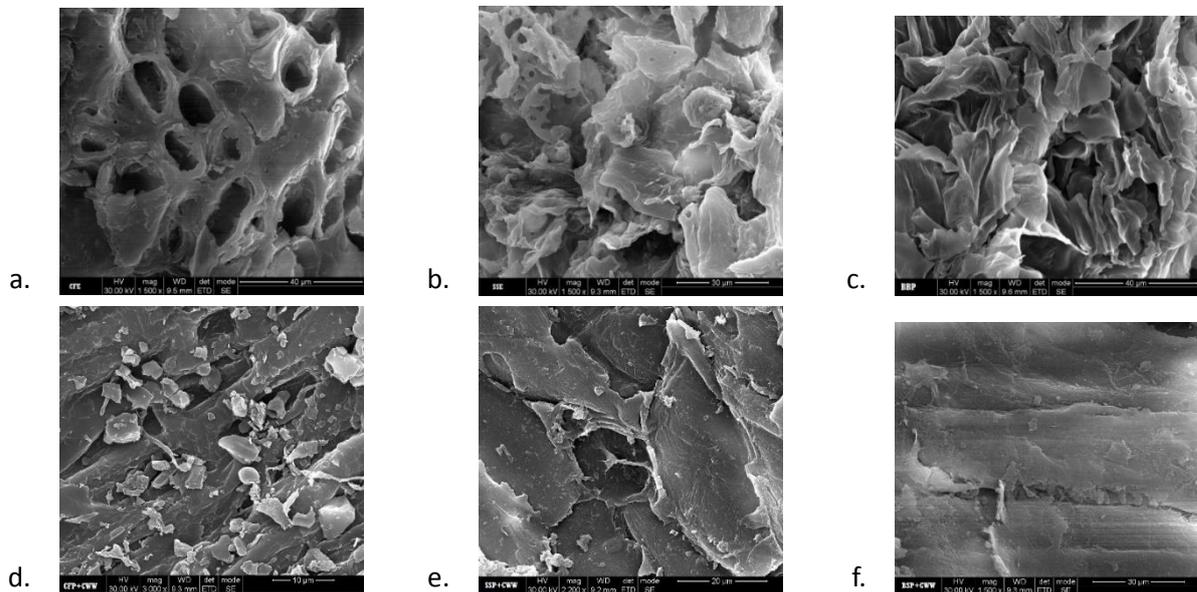
### 3.3 Effect of Adsorbent dose

With an increase in adsorbent dose, the percent reduction of COD increased, reached a maximum and then a slight decrease was observed (Fig. 4) which can be explained based on increase in number of pores/active sites as well as surface area with an increase in the amount of adsorbent (Rani devi *et al.*, 2008). The maximum percent reduction of COD was 99.2%, 92.7% and 81.0% (with  $\pm 0.1$  SD) using CFP, SSP and BSP respectively.

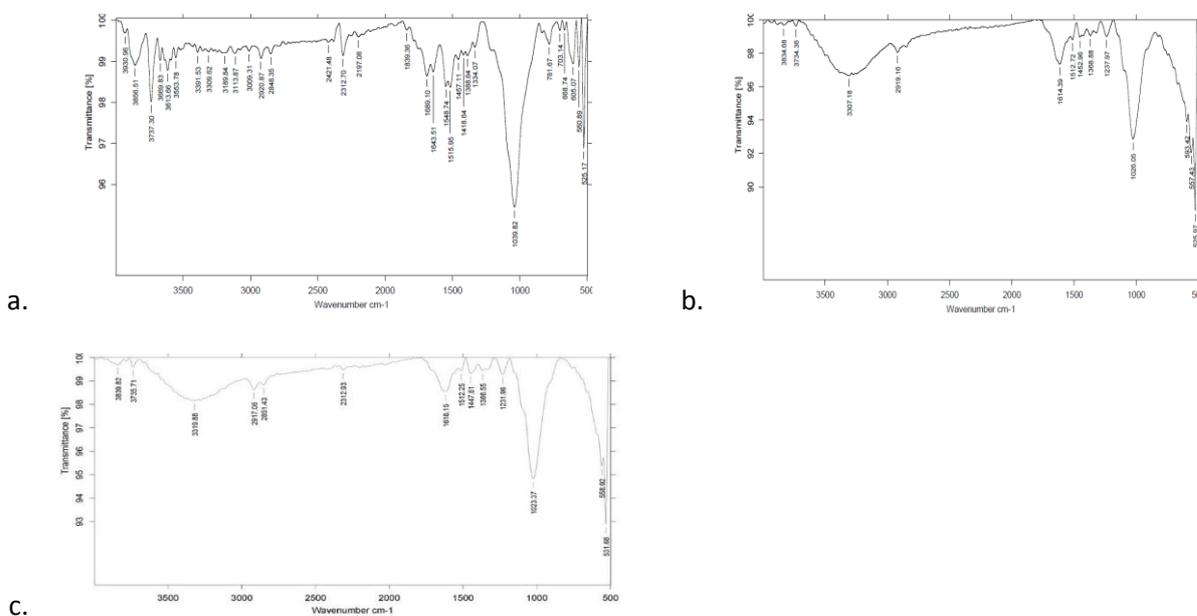
### 3.4 Effect of pH

The pH of solution can influence the degree of ionization of functional groups present on active sites of adsorbent. Though, equilibrium pH affect the biosorption capacity, its maintenance is difficult and hence, initial pH of the solution was kept constant as it has an influence on equilibrium pH. Moreover, sometimes, initial pH value of the solution has been shown to have more influence than the final pH (Waranusantigul *et al.*, 2003). In earlier studies by other workers, it is clear that there is no significant effect of lower pH on the reduction of COD (Rani Devi *et al.*, 2008) and hence, in the present case, study was carried out from pH 3 onwards with 3 hours as contact time. With an increase in pH, the percent of removal increased and reached maximum at neutral pH and then decreased (Fig. 5). Similar optimum pH at 7 and a fall in removal capacity afterwards was observed in the removal of COD from wastewater of coffee processing plant (Rani Devi *et al.*, 2008; Rani Devi, 2010) and sugar industrial effluent (Parande *et al.*, 2009). However,

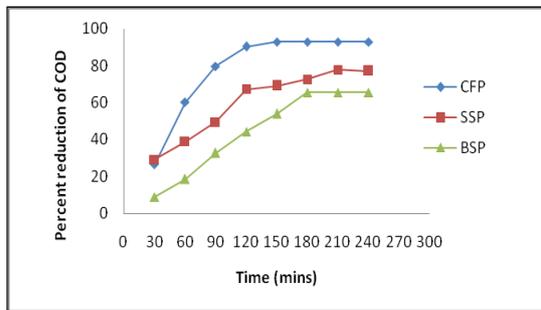
a slight enhancement in COD adsorption for pH values ranging from 7.5 to 8 was observed (Muftah H. El-Naas *et al.*, 2010) whereas, maximum COD removal was observed at pH 3.0–4.5 and 5.5 – 7.0 using activated coconut shell and industrial grade granular activated carbon respectively (Mahesh *et al.*, 1999). A decrease in percent of COD removal beyond pH 7 may be due to repulsion between the so generated negatively charged ions on adsorbent (by hydrolysis of phenol/alcohol and carboxylic acid) and in coffee waste water (by hydrolysis of various functional groups on biomolecules like pectin, fibers, fat, proteins, chlorogenic acid, lignins, tannins, humic acid). One more reasons for lower % removal at pH above 7 might be the higher solubility of oligosaccharides in alkaline solutions (Rothfos 1979; Treagust, 1994). Similarly, at lower pH, the presence larger number of H<sup>+</sup> ions neutralizes the negatively charged adsorbent surface and hence reduces COD removal by ion exchange (Tewari *et al.*, 2005).



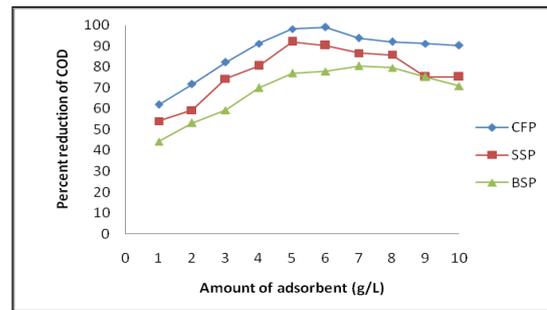
**Figure 1.** Scanning Electron Microscope images of untreated CFP, SSP and BSP (a, b and c) and after treatment effluent (d, e and f)



**Figure 2.** FT-IR spectra of three adsorbents (a) CFP and (b) SSP (c) BSP



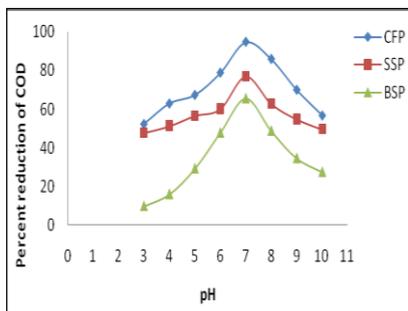
**Figure 3.** Effect of contact time on percent reduction of COD



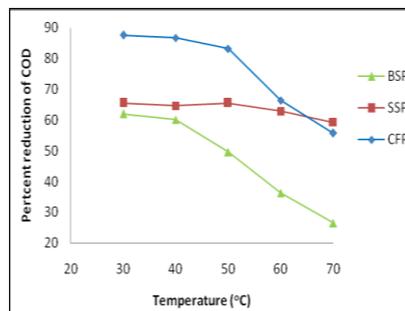
**Figure 4.** Effect of adsorbent dose on percent reduction of COD

**3.4 Effect of pH**

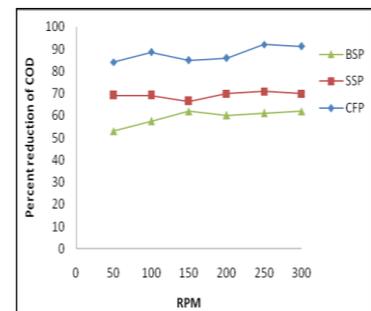
As the percent removal is maximum at pH 7, neutralization of coffee industry effluents has to be carried out prior to COD removal in adsorption methods because, recycling of the pulping water causes an increase in organic matter which ferments rapidly to produce organic acids leading to an increase in acidity (pH=4.3) (Adams and Dougan, 1987).



**Figure 5.** Effect of pH on %reduction of COD



**Figure 6.** Effect of temperature on %reduction of COD



**Figure 7.** Effect of agitation on %reduction of COD

**3.5 Effect of Temperature**

In general, preferable binding can be expected at higher temperatures for endothermic reactions, however, higher removal efficiencies are observed for exothermic reactions at lower temperatures. In the present case, a decrease in the adsorption percentage with an increase in temperature was observed with 3 hours as contact time. (Fig. 6). Though, this seems to be corroborates with the well known fact that sorption capacity is expected to decrease with an increase in solution temperature when sorption process is exothermic, the present case involves both physical and chemical adsorptions. Such simultaneous occurrence of both adsorptions was reported by (Ho and McKay 2003, Panday *et al.*, 1985).

Mostly, in the case of chemisorption, an increase in temperature increases adsorption capacity due to (i) increase in the kinetic energy of the adsorbate (ii) expansion of the pores within the adsorbent particles (Muftah H. El-Naas *et al.*, 2010) and (iii) swelling effect which enable to penetrate further (M. Doğan, M. Alkan, 2003). However, even in chemisorption, a decrease in removal capacity at an increased temperature can be explained based on (i) increased escaping tendency of already adsorbed molecules from adsorbent into the bulk solution due to increase in total energy of sorbate molecules at higher temperature (Sharma *et al.*, 1991) and (ii) dissolution of sorbing solution because, tannins and other phenol / alcohol compounds are soluble at high temperatures (Jin Dai and Russell J. Mumper, 2010). However, (Weber and Chakravorti, 1974) has stated that, since diffusion is an endothermic process, the rate of sorption will increase with increased solution temperature when intraparticle transport (pore diffusion) is the rate-limiting step. The results of this study on the effect of temperature indicate it is possible that this pseudo-

second-order model could act as a methodology for discriminating between sorption mechanisms on this basis.

Thermodynamic parameters ( $\Delta G$ ,  $\Delta H$  and  $\Delta S$  KJ mol<sup>-1</sup>) at 303 K are (-1.69, -34.17 and -0.10 KJ mol<sup>-1</sup> for BSP), (-5.58, -39.74 and -0.11 KJ mol<sup>-1</sup> for SSP) and (-1.71, -5.13 and -0.01 KJ mol<sup>-1</sup> for CFP). The activation energy  $E_a$  is calculated using Arrhenius equation. By plotting  $\log k^2$  versus  $1/T$  activation energy for dye adsorption is calculated as 19.65 kJ mol<sup>-1</sup>. As activation energy is lower than 40 kJ mol<sup>-1</sup>, it follows that the reaction is physical adsorption controlled process (Chowdhury *et al.*, 2009; Jaman *et al.*, 2009). Activation energy in the range of 17–21 KJ mol<sup>-1</sup> and 21–42 KJ mol<sup>-1</sup> indicates that the type of diffusion path of sorption process is film diffusion and pore diffusion respectively (Namsivayam and Sangeetha, 2006). Hence, in the present case, the diffusion is mostly pore diffusion.

### 3.6 Effect of agitation

In the case of BSP and SSP, no effect of agitation on the removal of COD was observed with 3 hours as contact time. But, a slight initial increase in adsorption capacity was noted in the case of CFP (Fig. 7), due to improved diffusion of sorbate towards sorbent and a decrease in film resistance to mass transfer surround the adsorbent particles with an initial increase in agitation. A slight oscillatory nature in adsorption capacity at higher agitation speeds can be explained based on the competition between improved diffusion and desorption of sorbate due to centrifugal force / mass transfer resistance contributes at higher agitations (Crini *et al.*, 2007).

### 3.7 Adsorption Isotherm studies

Langmuir, Freundlich and D-R isotherms were employed at the pre-determined equilibrium conditions in order to analyze the interaction of adsorbate on the surface of each adsorbent (Table.2). Based on the values of  $b/R_L$ , Langmuir equation is favorable ( $0 < R_L < 1$ ) or unfavorable ( $R_L > 1$ ). Similarly, the value of sorption intensity ( $1 < n < 10$ ) indicates that sorption is Freundlich favorable. As the  $R_L$  values in the present case are in the range of 0.011 - 0.016 and 'n' values are less than 10, the three adsorbents (CFP, SSP and BSP), Langmuir as well as Freundlich isotherms are favorable for adsorption process, whereas, based on the values of regression coefficients, it is clear that Langmuir isotherm fit relatively better than the Freundlich isotherm. Hence, both mono and multilayer adsorptions occur in the present studies with preference to monolayer adsorption. A similar observation on data confirmation to combination of mechanisms – Freundlich and Langmuir isotherms was reported in reduction of COD in sugar industrial effluent (Parande *et al.*, 2009) and removal of dyes (Ho and McKay *et al.*, 2003).

**Table 2.** Adsorption isotherm parameters of three adsorbents, CFP, SSP and BSP

| Isotherm   | Parameter                                | CFP     | SSP     | BSP    |
|------------|--|---------|---------|--------|
| Langmuir   | <b>b (l mg<sup>-1</sup>)</b>             | 0.044   | 0.065   | 0.060  |
|            | <b>q<sub>m</sub> (mg g<sup>-1</sup>)</b> | 27.78   | 28.57   | 62.50  |
|            | <b>R<sup>2</sup></b>                     | 0.685   | 0.909   | 0.996  |
| Freundlich | <b>n</b>                                 | 1.252   | 3.559   | 9.091  |
|            | <b>k<sub>F</sub></b>                     | 27.91   | 8.48    | 4.45   |
|            | <b>R<sup>2</sup></b>                     | 0.420   | 0.766   | 0.861  |
| D-R        | <b>q<sub>D</sub> (mg g<sup>-1</sup>)</b> | 1644.21 | 1203.51 | 582.39 |
|            | <b>B<sub>D</sub></b>                     | -0.137  | -0.934  | -3.316 |
|            | <b>R<sup>2</sup></b>                     | 0.524   | 0.877   | 0.908  |

### 3.8 Adsorption kinetics

The mechanism of action of adsorbate-adsorbent interactions has been analysed in terms of pseudo first order (Lagergren *et al.*, 1898) and pseudo second order sorption (Ho and McKay, 1999) equations.

In the case of linearized pseudo-first order equation, it is observed that the results deviate significantly from a straight line, which indicates that the pore diffusion is not the sole rate controlling step. Moreover, the deviation, the values of  $R^2$  from unity is a measure of the incompatibility of each respective case. In

the present case, the first order kinetics show poor correlation coefficients. On the other hand, when  $t/q$  was plotted against  $t$ , according to the linearized pseudo-second order model, a clear linear relationship was observed for all cases. The data were fitted perfectly well by straight lines, with  $R^2$  values in the range of 0.993 to 0.999. This proves that the adsorption kinetics is more accurately described by pseudo-second order model. Higher values of regression coefficient  $R^2$  were observed for second order kinetics compared to first order. In addition, the  $q_e$  values (calculated) from pseudo-second-order model agree perfectly with the experimental  $q_e$  values (Table.3). Pseudo-second order kinetic data reveal the mechanism of adsorbate and adsorbent interaction as chemical coordination in addition to physical attachment. Moreover, the correlation coefficients ( $R^2$ ) for the pseudo-second order kinetic model are much greater than the intra-particle diffusion coefficients which strongly suggest a chemical reaction mechanism (Tables. 3 and 4).

**Table 3.** Adsorption kinetic parameters, of adsorbents, CFP, SSP and BSP

| Adsorbent<br>Sample     | $q_e$ $\text{mgg}^{-1}$ (exp) | First order                   |                       |       | Second order                  |  |       |
|-------------------------|-------------------------------|-------------------------------|-----------------------|-------|-------------------------------|--|-------|
|                         |                               | $q_e$ $\text{mgg}^{-1}$ (cal) | $K$ $\text{min}^{-1}$ | $R^2$ | $q_e$ $\text{mgg}^{-1}$ (cal) | $H$ ( $\text{mgg}^{-1}\text{min}^{-1}$ ) | $R^2$ |
| CFP/5g $\text{l}^{-1}$  | 23.20                         | 2.69                          | 0.009                 | 0.810 | 28.6                          | 1.25                                     | 0.997 |
| CFP/10g $\text{l}^{-1}$ | 35.00                         | 1.76                          | 0.005                 | 0.800 | 37.0                          | 5.71                                     | 0.997 |
| CFP/20g $\text{l}^{-1}$ | 63.45                         | 1.87                          | 0.012                 | 0.995 | 66.7                          | 18.52                                    | 0.999 |
| SSP/5g $\text{l}^{-1}$  | 25.00                         | 1.92                          | 0.005                 | 0.500 | 28.6                          | 1.99                                     | 0.993 |
| SSP/10g $\text{l}^{-1}$ | 47.30                         | 2.05                          | 0.016                 | 0.921 | 50.0                          | 11.11                                    | 0.999 |
| SSP/20g $\text{l}^{-1}$ | 67.30                         | 2.09                          | 0.018                 | 0.977 | 71.4                          | 23.36                                    | 0.999 |
| BSP/5g $\text{l}^{-1}$  | 45.60                         | 2.01                          | 0.012                 | 0.750 | 47.6                          | 9.27                                     | 0.998 |
| BSP/10g $\text{l}^{-1}$ | 48.20                         | 1.44                          | 0.005                 | 0.899 | 50.0                          | 16.49                                    | 0.999 |
| BSP/20g $\text{l}^{-1}$ | 57.20                         | 1.48                          | 0.007                 | 0.750 | 58.8                          | 23.81                                    | 0.999 |

Similarly, in the reduction of COD in refinery wastewater on date-pit activated carbon, kinetic data were best fitted by the pseudo-second order model, and the equilibrium data followed the Langmuir monolayer isotherm (Muftah H. El-Naas *et al.*, 2010).

### 3.9 Boyd, Weber's and Elovich models in support of adsorption mechanism

To investigate the mechanism of adsorption, the kinetic data obtained for the adsorption were evaluated using Boyd's kinetic model (Boyd *et al.*, 1947) and Weber's intraparticle model (Weber and Morris, 1963). Though Boyd plots were reasonably linear ( $R^2 = 0.902, 0.949$  and  $0.938$ ), but have intercepts, hence, internal diffusion may not be the sole rate controlling step in the adsorption process (Poots *et al.*, 1976).

**Table 4.** Weber's intraparticle, Boyd and Elovich's adsorption kinetic parameters, of adsorbents, CFP, SSP and BSP

| Adsorbent                                       | CFP    | SSP    | BSP     |
|---|--------|--------|---------|
| <b>Weber's Intraparticle diffusion</b>          |        |        |         |
| $k_d$   | 174.4  | 147.9  | 180.3   |
| $q$   | 270.90 | 48.65  | -759.80 |
| $R^2$   | 0.789  | 0.946  | 0.967   |
| <b>Boyd model</b>                               |        |        |         |
| $R^2$   | 0.938  | 0.949  | 0.902   |
| <b>Elovich's kinetics</b>                       |        |        |         |
| $\alpha$ ( $\text{mg g}^{-1} \text{min}^{-1}$ ) | 92.48  | 66.55  | 32.45   |
| $1/\beta$ ( $\text{mg g}^{-1}$ )                | 909.80 | 728.10 | 877.80  |
| $R^2$   | 0.896  | 0.958  | 0.957   |

Multi-linearity in Weber's plots indicates the involvement of more than one kinetic stage. Mainly three

stages were observed viz., diffusion of sorbate to sorbent (fast stage), intra particle diffusion (delayed) and diffusion through small pores (establishment of equilibrium). This trend is similar to that observed by Chen and Liu, (2012). However, in the reduction of COD in refinery wastewater on date-pit activated carbon, modeling of the controlling mechanisms indicated that both intrinsic kinetics and mass transfer contributed to controlling the adsorption process. Mass transfer seemed to be the dominant mechanism at low COD content, while intrinsic kinetics dominates at high concentrations (Muftah H. El-Naas *et al.*, 2010).

The Elovich equation was also satisfied with reasonable linearity supporting involvement of chemisorptions as the one of the mechanism and  $\alpha$  (initial sorption rate) as well as  $\beta$  (related with extent of surface coverage and activation energy for chemisorption) are tabulated (Table.4).

The comparison of the percent of COD removal in the present study with earlier studies show that Casuarina fruit powder and Sorghum stem powder work extremely well, whereas, Banana stem powder is reasonably acceptable.

| Adsorbent                     | %COD reduced | Reference                      |
|-------------------------------|--------------|--------------------------------|
| Avacado peel carbon           | 98.2         | Rani Devi <i>et al.</i> , 2008 |
| Tamarind nut carbon           | 70.0         | Parande <i>et al.</i> , 2009   |
| Dates nut carbon              | 72.2         | Parande <i>et al.</i> , 2009   |
| Metakaolin                    | 75.3         | Parande <i>et al.</i> , 2009   |
| Crushed burnt coal            | 85.0         | Debabrata, 2011                |
| Fly ash                       | 25.0         | Lakdawala & Lakdawala, 2012    |
| Coconut coir activated carbon | 80.0         | Kulkarni, 2013                 |
| Casuarina fruit powder        | 99.2         | Current study                  |
| Sorghum stem powder           | 92.7         | Current study                  |
| Banana stem powder            | 81.0         | Current study                  |

#### 4. Conclusion

The maximum adsorption efficiencies of the three adsorbents for COD removal from coffee waste water were 99.2%, 92.7% and 81.0% using CFP, SSP and BSP respectively. The order of adsorption efficiency was CFP>SSP>BSP can be correlated with their microporous structures, where, CFP sample surface consists of cavities indicating high possibility of adsorption. The proposed involvement of chemisorption was correlated with the possible interactions between functional groups present on biosorbents and sorbates (contributing to COD in coffee waste water). Langmuir isotherm fit relatively better than the Freundlich isotherm and the adsorption data showed good agreement with the pseudo-second-order kinetic model. But at the same time, the adsorption capacity decreased with an increase in temperature, indicating endothermic nature of adsorption. Based on the observations, both mono and multilayer adsorptions occur in the present studies with preference to monolayer adsorption. Considering the fact that these three adsorbents are abundantly available, the efficient reduction of COD from wastewater could be achieved by utilizing these adsorbents industrially.

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